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3-(4-Fluorobenzyl)-1*H*-isochromene-1-thione

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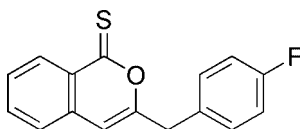
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Key indicators: single-crystal X-ray study; $T = 294$ K; mean $\sigma(\text{C}-\text{C}) = 0.003$ Å; R factor = 0.046; wR factor = 0.150; data-to-parameter ratio = 18.5.

In the molecule of the title compound, $\text{C}_{16}\text{H}_{11}\text{FOS}$, the benzene ring is oriented at a dihedral angle of 89.68 (3) $^\circ$ with respect to the planar [maximum deviation 0.009 (2) Å] isocoumarin ring system. An intramolecular $\text{C}-\text{H}\cdots\text{S}$ interaction results in the formation of a planar five-membered ring. In the crystal structure, intermolecular $\text{C}-\text{H}\cdots\text{O}$ hydrogen bonds link the molecules into chains parallel to the c axis. A $\pi-\pi$ contact between the isocoumarin rings [centroid-centroid distance = 3.818 (3) Å] may further stabilize the structure.

Related literature

For general background, see: Barry (1964); Sturtz *et al.* (2002); Rossi *et al.* (2003); Powers *et al.* (2002); Thomas & Jens (1999). For a related structure, see: Abid *et al.* (2006). For bond-length data, see: Allen *et al.* (1987).



Experimental

Crystal data

$\text{C}_{16}\text{H}_{11}\text{FOS}$ $b = 17.9516$ (11) Å
 $M_r = 270.31$ $c = 8.4481$ (5) Å
 Monoclinic, $P2_1/c$ $\beta = 95.026$ (1) $^\circ$
 $a = 8.7346$ (6) Å $V = 1319.57$ (14) Å 3

$Z = 4$
 Mo $K\alpha$ radiation
 $\mu = 0.25$ mm^{-1}

$T = 294$ (2) K
 $0.30 \times 0.25 \times 0.20$ mm

Data collection

Bruker SMART CCD area-detector diffractometer 7856 measured reflections
 3188 independent reflections
 Absorption correction: multi-scan (SADABS; Bruker, 2001) 2655 reflections with $I > 2\sigma(I)$
 $R_{\text{int}} = 0.016$
 $T_{\text{min}} = 0.805$, $T_{\text{max}} = 0.952$

Refinement

$R[F^2 > 2\sigma(F^2)] = 0.046$ 172 parameters
 $wR(F^2) = 0.150$ H-atom parameters constrained
 $S = 1.03$ $\Delta\rho_{\text{max}} = 0.34$ $\text{e} \text{ \AA}^{-3}$
 3188 reflections $\Delta\rho_{\text{min}} = -0.30$ $\text{e} \text{ \AA}^{-3}$

Table 1
 Hydrogen-bond geometry (Å, $^\circ$).

$D-\text{H}\cdots A$	$D-\text{H}$	$\text{H}\cdots A$	$D\cdots A$	$D-\text{H}\cdots A$
$\text{C1}-\text{H1A}\cdots\text{S1}$	0.93	2.78	3.142 (2)	105
$\text{C3}-\text{H3A}\cdots\text{O1}^i$	0.93	2.60	3.529 (2)	178

Symmetry code: (i) $x, y, z - 1$.

Data collection: SMART (Bruker, 2001); cell refinement: SAINT (Bruker, 2002); data reduction: SAINT; program(s) used to solve structure: SHELXS97 (Sheldrick, 2008); program(s) used to refine structure: SHELXL97 (Sheldrick, 2008); molecular graphics: ORTEP-3 for Windows (Farrugia, 1997) and PLATON (Spek, 2003); software used to prepare material for publication: SHELXTL (Sheldrick, 2008) and PLATON.

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Supplementary data and figures for this paper are available from the IUCr electronic archives (Reference: HK2617).

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supplementary materials

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3-(4-Fluorobenzyl)-1*H*-isochromene-1-thione

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Comment

The isocoumarin nucleus is an abundant structural motif in natural products (Barry, 1964). Many constituents of the steadily growing class of known isocoumarins exhibit valuable biological properties such as antifungal (Sturtz *et al.*, 2002), antitumor or cytotoxic, anti-inflammatory, anti-allergic (Rossi *et al.*, 2003) and enzyme inhibitory (Powers *et al.*, 2002) activities. Naturally occurring halo-isocoumarins and their halogeno-3,4-dihydroisocoumarin derivatives are very rare. However, a few examples of naturally occurring chlorine containing isocoumarins are known (Thomas & Jens, 1999). In view of the importance of this class of compounds, the title compound, an isocoumarine derivative containing 4-fluorobenzyl substituent has been synthesized, and we report herein its crystal structure.

In the molecule of the title compound (Fig. 1), the bond lengths (Allen *et al.*, 1987) and angles are within normal ranges, and comparable with the corresponding values in 3-(2-chlorobenzyl)isocoumarin (Abid *et al.*, 2006). Rings A (C1-C6), B (O1/C5-C9) and C (C11-C16) are, of course, planar and the dihedral angles between them are A/B = 0.29 (3)°, A/C = 89.77 (4)° and B/C = 89.53 (3)°. The intramolecular C-H...S interaction (Table 1) results in the formation of a planar five-membered ring D (S1/C1/C6/C7/H1A).

In the crystal structure, intermolecular C-H...O hydrogen bonds (Table 1) link the molecules into chains parallel to the *c*-axis (Fig. 2), in which they may be effective in the stabilization of the structure. The π - π contact between the isocoumarine rings, Cg1—Cg2ⁱ [symmetry code: (i) -*x*, -*y*, 1 - *z*, where Cg1 and Cg2 are centroids of the rings A (C1-C6) and B (O1/C5-C9), respectively] may further stabilize the structure, with centroid-centroid distance of 3.818 (3) Å.

Experimental

As shown in Scheme 2, the title compound was synthesized by refluxing 3-(4-fluorobenzyl)-1*H*-isochromen-1-one (0.5 g, 1.8 mmol) with Lawesson's reagent (0.89 g, 2.2 mmol) in dry toluene for 4 h. Pure thioisocoumarin was obtained by recrystallization in methanol (yield; 90%, m.p. 665-667 K).

Refinement

H atoms were positioned geometrically, with C-H = 0.93 and 0.97 Å for aromatic and methylene H, respectively, and constrained to ride on their parent atoms, with $U_{\text{iso}}(\text{H}) = 1.2U_{\text{eq}}(\text{C})$.

Figures

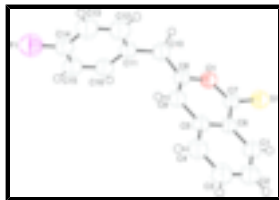


Fig. 1. The molecular structure of the title molecule, with the atom-numbering scheme.

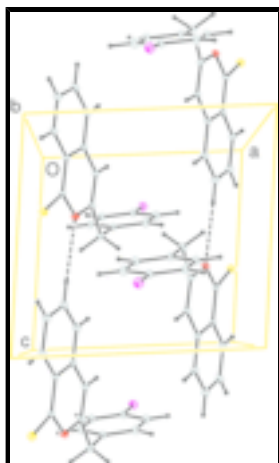


Fig. 2. A partial packing diagram of the title compound. Hydrogen bonds are shown as dashed lines.



Fig. 3. The formation of the title compound.

3-(4-Fluorobenzyl)-1*H*-isochromene-1-thione

Crystal data

$C_{16}H_{11}FOS$

$M_r = 270.31$

Monoclinic, $P2_1/c$

Hall symbol: $-P\ 2ybc$

$a = 8.7346\ (6)\ \text{\AA}$

$b = 17.9516\ (11)\ \text{\AA}$

$c = 8.4481\ (5)\ \text{\AA}$

$\beta = 95.026\ (1)^\circ$

$V = 1319.57\ (14)\ \text{\AA}^3$

$Z = 4$

$F_{000} = 560$

$D_x = 1.361\ \text{Mg m}^{-3}$

Melting point: $392(2)\ \text{K}$

Mo $K\alpha$ radiation

$\lambda = 0.71073\ \text{\AA}$

Cell parameters from 1423 reflections

$\theta = 4.2\text{--}25.4^\circ$

$\mu = 0.25\ \text{mm}^{-1}$

$T = 294\ \text{K}$

Block, yellow

$0.30 \times 0.25 \times 0.20\ \text{mm}$

Data collection

Bruker SMART CCD area-detector diffractometer

Radiation source: fine-focus sealed tube

Monochromator: graphite

$T = 294\ \text{K}$

3188 independent reflections

2655 reflections with $I > 2\sigma(I)$

$R_{\text{int}} = 0.016$

$\theta_{\text{max}} = 28.3^\circ$

φ and ω scans	$\theta_{\min} = 2.6^\circ$
Absorption correction: multi-scan (SADABS; Bruker, 2001)	$h = -11 \rightarrow 11$
$T_{\min} = 0.805$, $T_{\max} = 0.952$	$k = -23 \rightarrow 22$
7856 measured reflections	$l = -8 \rightarrow 11$

Refinement

Refinement on F^2	Secondary atom site location: difference Fourier map
Least-squares matrix: full	Hydrogen site location: inferred from neighbouring sites
$R[F^2 > 2\sigma(F^2)] = 0.046$	H-atom parameters constrained
$wR(F^2) = 0.150$	$w = 1/[\sigma^2(F_o^2) + (0.0856P)^2 + 0.2983P]$
$S = 1.03$	where $P = (F_o^2 + 2F_c^2)/3$
3188 reflections	$(\Delta/\sigma)_{\max} < 0.001$
172 parameters	$\Delta\rho_{\max} = 0.34 \text{ e } \text{\AA}^{-3}$
Primary atom site location: structure-invariant direct methods	$\Delta\rho_{\min} = -0.29 \text{ e } \text{\AA}^{-3}$
	Extinction correction: none

Special details

Geometry. All e.s.d.'s (except the e.s.d. in the dihedral angle between two l.s. planes) are estimated using the full covariance matrix. The cell e.s.d.'s are taken into account individually in the estimation of e.s.d.'s in distances, angles and torsion angles; correlations between e.s.d.'s in cell parameters are only used when they are defined by crystal symmetry. An approximate (isotropic) treatment of cell e.s.d.'s is used for estimating e.s.d.'s involving l.s. planes.

Refinement. Refinement of F^2 against ALL reflections. The weighted R -factor wR and goodness of fit S are based on F^2 , conventional R -factors R are based on F , with F set to zero for negative F^2 . The threshold expression of $F^2 > \sigma(F^2)$ is used only for calculating R -factors(gt) *etc.* and is not relevant to the choice of reflections for refinement. R -factors based on F^2 are statistically about twice as large as those based on F , and R -factors based on ALL data will be even larger.

Fractional atomic coordinates and isotropic or equivalent isotropic displacement parameters (\AA^2)

	x	y	z	$U_{\text{iso}}^*/U_{\text{eq}}$
S1	0.07012 (6)	0.65948 (3)	0.36335 (6)	0.0646 (2)
O1	0.20272 (15)	0.53300 (6)	0.37410 (13)	0.0506 (3)
F1	0.51024 (16)	0.13358 (7)	0.29180 (19)	0.0845 (4)
C1	0.0938 (2)	0.63191 (10)	-0.0010 (2)	0.0563 (4)
H1A	0.0472	0.6745	0.0350	0.068*
C2	0.1045 (3)	0.62201 (12)	-0.1609 (2)	0.0665 (5)
H2A	0.0656	0.6581	-0.2326	0.080*
C3	0.1727 (3)	0.55866 (12)	-0.2156 (2)	0.0706 (6)
H3A	0.1785	0.5521	-0.3241	0.085*
C4	0.2320 (3)	0.50543 (11)	-0.1113 (2)	0.0656 (5)
H4A	0.2784	0.4633	-0.1496	0.079*
C5	0.2234 (2)	0.51390 (9)	0.05279 (18)	0.0483 (4)
C6	0.15271 (18)	0.57814 (8)	0.10781 (17)	0.0439 (3)

supplementary materials

C7	0.14344 (18)	0.58791 (9)	0.27669 (18)	0.0446 (3)
C8	0.2707 (2)	0.46929 (8)	0.32005 (19)	0.0482 (4)
C9	0.2829 (2)	0.45929 (9)	0.1661 (2)	0.0523 (4)
H9A	0.3304	0.4166	0.1313	0.063*
C10	0.3182 (3)	0.41971 (10)	0.4583 (2)	0.0615 (5)
H10A	0.2320	0.4141	0.5222	0.074*
H10B	0.4006	0.4438	0.5237	0.074*
C11	0.3716 (2)	0.34333 (9)	0.4123 (2)	0.0515 (4)
C12	0.5239 (3)	0.32831 (12)	0.3957 (3)	0.0741 (6)
H12A	0.5961	0.3662	0.4129	0.089*
C13	0.5718 (2)	0.25794 (14)	0.3540 (3)	0.0799 (7)
H13A	0.6748	0.2483	0.3421	0.096*
C14	0.4648 (2)	0.20356 (10)	0.3309 (2)	0.0585 (4)
C15	0.3136 (2)	0.21543 (11)	0.3461 (3)	0.0671 (5)
H15A	0.2425	0.1771	0.3292	0.081*
C16	0.2677 (2)	0.28622 (11)	0.3875 (3)	0.0644 (5)
H16A	0.1644	0.2952	0.3987	0.077*

Atomic displacement parameters (\AA^2)

	U^{11}	U^{22}	U^{33}	U^{12}	U^{13}	U^{23}
S1	0.0761 (4)	0.0604 (3)	0.0583 (3)	0.0200 (2)	0.0112 (2)	-0.0054 (2)
O1	0.0702 (8)	0.0431 (6)	0.0390 (5)	0.0077 (5)	0.0071 (5)	-0.0006 (4)
F1	0.0826 (9)	0.0557 (7)	0.1132 (11)	0.0188 (6)	-0.0030 (8)	-0.0146 (7)
C1	0.0629 (11)	0.0529 (9)	0.0526 (9)	0.0022 (8)	0.0028 (7)	0.0071 (7)
C2	0.0840 (14)	0.0661 (11)	0.0485 (9)	-0.0040 (10)	-0.0003 (9)	0.0157 (9)
C3	0.1044 (17)	0.0684 (12)	0.0396 (8)	-0.0116 (11)	0.0106 (9)	0.0040 (8)
C4	0.1026 (16)	0.0520 (9)	0.0441 (9)	-0.0023 (10)	0.0173 (9)	-0.0040 (7)
C5	0.0635 (10)	0.0415 (7)	0.0408 (7)	-0.0063 (7)	0.0093 (7)	-0.0009 (6)
C6	0.0473 (8)	0.0431 (7)	0.0412 (7)	-0.0052 (6)	0.0042 (6)	0.0012 (6)
C7	0.0462 (8)	0.0438 (7)	0.0437 (7)	-0.0001 (6)	0.0043 (6)	0.0000 (6)
C8	0.0623 (10)	0.0364 (7)	0.0460 (8)	0.0016 (6)	0.0053 (7)	-0.0004 (6)
C9	0.0713 (11)	0.0383 (7)	0.0484 (8)	0.0032 (7)	0.0110 (7)	-0.0020 (6)
C10	0.0910 (14)	0.0465 (9)	0.0466 (9)	0.0090 (9)	0.0037 (8)	0.0035 (7)
C11	0.0653 (10)	0.0424 (8)	0.0464 (8)	0.0041 (7)	0.0019 (7)	0.0069 (6)
C12	0.0598 (12)	0.0560 (11)	0.1061 (18)	-0.0107 (9)	0.0047 (11)	-0.0067 (11)
C13	0.0506 (11)	0.0685 (13)	0.121 (2)	0.0047 (9)	0.0111 (12)	-0.0073 (13)
C14	0.0621 (11)	0.0453 (8)	0.0666 (11)	0.0093 (7)	-0.0027 (8)	-0.0016 (8)
C15	0.0561 (11)	0.0462 (9)	0.0976 (15)	-0.0034 (8)	-0.0018 (10)	-0.0022 (9)
C16	0.0508 (10)	0.0538 (10)	0.0891 (14)	0.0065 (8)	0.0080 (9)	0.0029 (9)

Geometric parameters (\AA , $^\circ$)

C1—C2	1.373 (3)	C8—C10	1.498 (2)
C1—C6	1.400 (2)	C9—H9A	0.9300
C1—H1A	0.9300	C10—C11	1.510 (2)
C2—C3	1.382 (3)	C10—H10A	0.9700
C2—H2A	0.9300	C10—H10B	0.9700
C3—C4	1.370 (3)	C11—C16	1.373 (3)

C3—H3A	0.9300	C11—C12	1.376 (3)
C4—C5	1.403 (2)	C12—C13	1.386 (3)
C4—H4A	0.9300	C12—H12A	0.9300
C5—C6	1.406 (2)	C13—C14	1.353 (3)
C5—C9	1.435 (2)	C13—H13A	0.9300
C6—C7	1.447 (2)	C14—C15	1.355 (3)
C7—O1	1.3573 (19)	C14—F1	1.367 (2)
C7—S1	1.6367 (16)	C15—C16	1.387 (3)
C8—C9	1.327 (2)	C15—H15A	0.9300
C8—O1	1.3843 (18)	C16—H16A	0.9300
C2—C1—C6	120.26 (18)	C5—C9—H9A	119.8
C2—C1—H1A	119.9	C8—C10—C11	114.22 (15)
C6—C1—H1A	119.9	C8—C10—H10A	108.7
C1—C2—C3	120.26 (18)	C11—C10—H10A	108.7
C1—C2—H2A	119.9	C8—C10—H10B	108.7
C3—C2—H2A	119.9	C11—C10—H10B	108.7
C4—C3—C2	120.55 (17)	H10A—C10—H10B	107.6
C4—C3—H3A	119.7	C16—C11—C12	118.03 (17)
C2—C3—H3A	119.7	C16—C11—C10	120.16 (18)
C3—C4—C5	120.66 (18)	C12—C11—C10	121.81 (18)
C3—C4—H4A	119.7	C11—C12—C13	121.35 (19)
C5—C4—H4A	119.7	C11—C12—H12A	119.3
C4—C5—C6	118.58 (15)	C13—C12—H12A	119.3
C4—C5—C9	122.49 (16)	C14—C13—C12	118.3 (2)
C6—C5—C9	118.92 (14)	C14—C13—H13A	120.8
C1—C6—C5	119.68 (15)	C12—C13—H13A	120.8
C1—C6—C7	121.00 (15)	C13—C14—C15	122.61 (18)
C5—C6—C7	119.31 (14)	C13—C14—F1	119.12 (18)
O1—C7—C6	117.29 (13)	C15—C14—F1	118.27 (18)
O1—C7—S1	116.25 (11)	C14—C15—C16	118.29 (18)
C6—C7—S1	126.45 (12)	C14—C15—H15A	120.9
C9—C8—O1	120.60 (14)	C16—C15—H15A	120.9
C9—C8—C10	130.03 (16)	C11—C16—C15	121.39 (18)
O1—C8—C10	109.36 (13)	C11—C16—H16A	119.3
C8—C9—C5	120.38 (15)	C15—C16—H16A	119.3
C8—C9—H9A	119.8	C7—O1—C8	123.49 (12)

Hydrogen-bond geometry (Å, °)

<i>D</i> —H... <i>A</i>	<i>D</i> —H	H... <i>A</i>	<i>D</i> ... <i>A</i>	<i>D</i> —H... <i>A</i>
C1—H1A...S1	0.93	2.78	3.142 (2)	105
C3—H3A...O1 ⁱ	0.93	2.60	3.529 (2)	178

Symmetry codes: (i) *x*, *y*, *z*−1.

Fig. 1

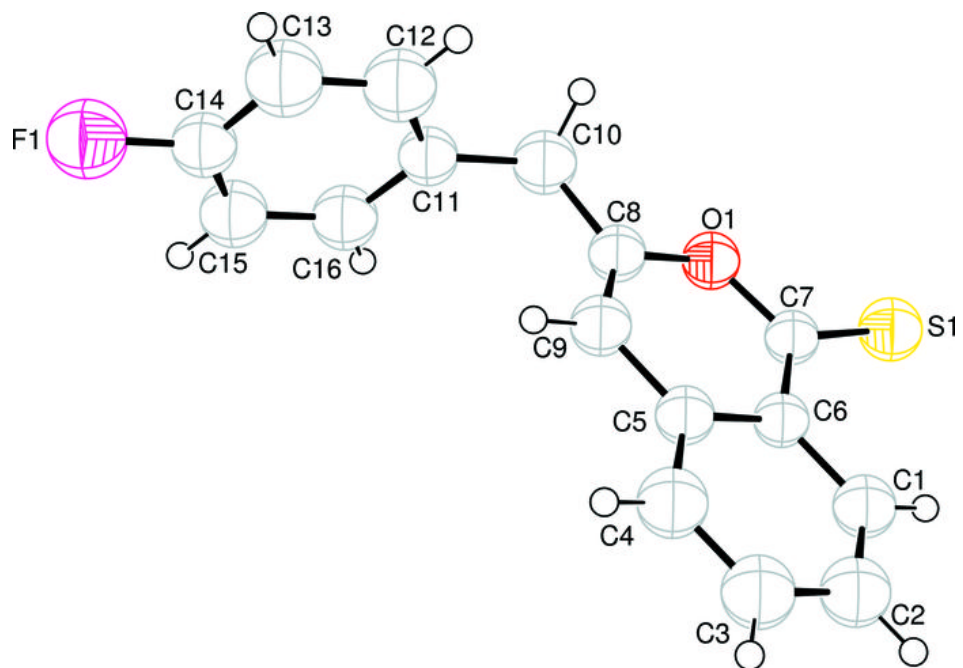


Fig. 2

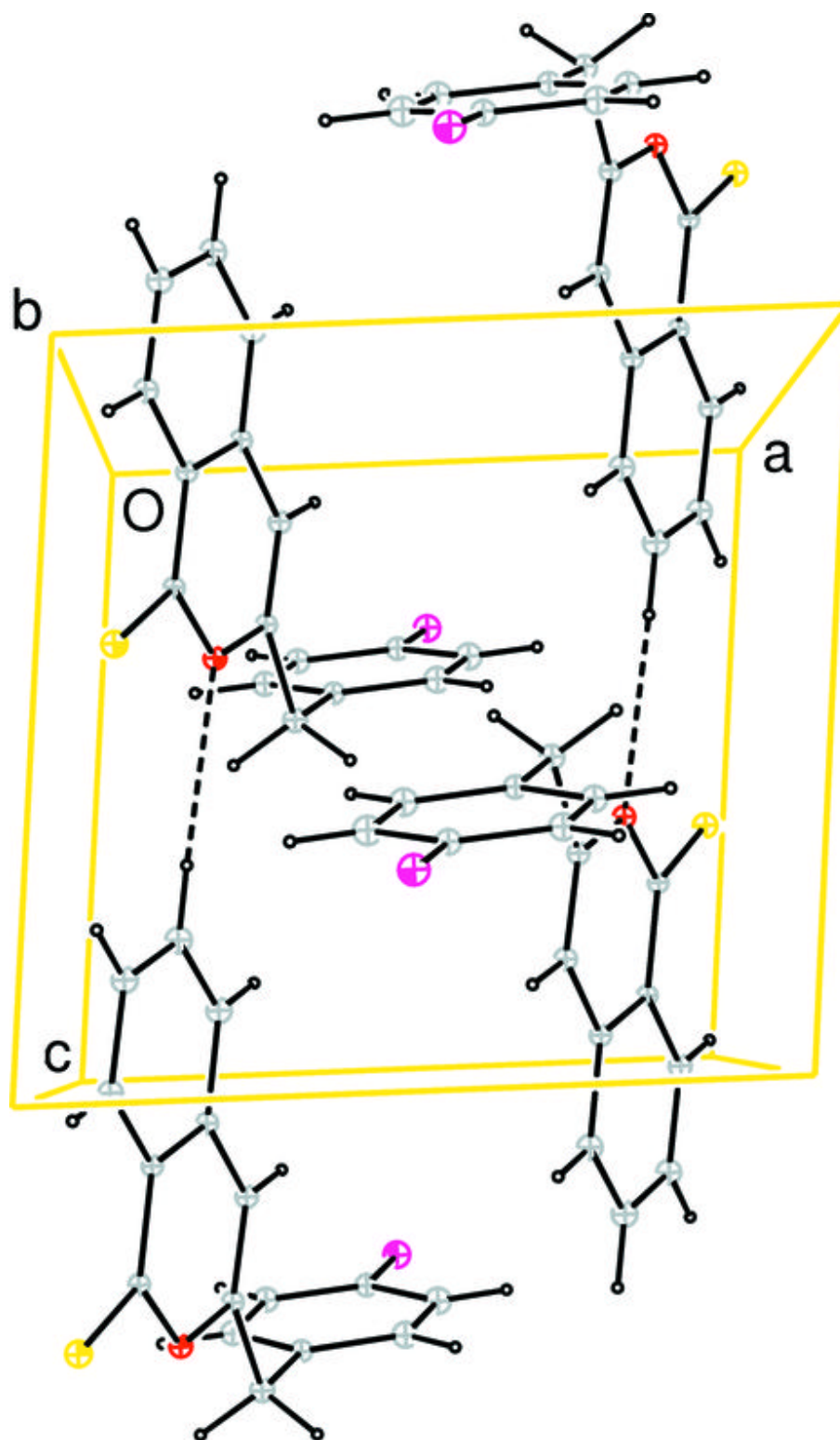


Fig. 3

